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What intermolecular forces are present in ch3br

What types of intermolecular forces are present in ch3br. What intermolecular forces are present in ch3cl. What intermolecular forces are present in ch3f. What intermolecular forces are present in ch3nh2. Ch3 intermolecular forces. What intermolecular forces are present in ch4.

D3 Dimension Technology Pte. Ltd. Address: 152 Beach Road #11-05 The Gateway Singapore (189721) ### The intermolecular forces in CH3Br are primarily due to dipole-dipole attraction and London dispersion forces, but the strongest force present is dipole-dipole attraction. ### Intermolecular forces play a crucial role in determining the physical properties of compounds. Among the given compounds - carbon dioxide (CO2), methanol (CH3OH), rubidium fluoride (RbF), and methyl bromide (CH3Br) - their boiling points can be ranked based on the strengths of these intermolecular forces. - **Carbon Dioxide (CO2)**: The weakest type of intermolecular force, London dispersion forces, are present in CO2. Given its small size compared to the other compounds, it has relatively weak intermolecular forces. - **Methyl Bromide (CH3Br)** and **Rubidium Fluoride (RbF)**: These two compounds exhibit dipole-dipole interactions and dispersion forces, respectively. While CH3Br shows dipole-dipole interactions due to its polar nature, RbF's ionic character leads to the presence of both dispersion and ionic forces. - **Methanol (CH3OH)**: Methanol boasts the strongest intermolecular forces among these compounds. The hydrogen bonding present in methanol significantly contributes to its high boiling point. Its intermolecular forces are stronger than those of methyl bromide, making CH3OH have a higher boiling point than CH3Br. Considering these factors, the order from lowest to highest boiling points is: CO2 < RbF < CH3Br < CH3OH. This ranking reflects the relative strengths of their intermolecular forces. Ionic forces, in fact, can be a dipole.